



# Groups and Periods on the Periodic Table

1. Write the group numbers in the boxes along the top of the periodic table.
2. Write the period numbers in the boxes down the side of the periodic table.

[illegible]

Elements in the same group have similar properties.

Lithium (Li) and potassium (K) are metals. They are solids at room temperature, have low melting points compared to other metals and are very reactive.

3. Which other elements have similar properties to lithium and potassium?

Tick **two** boxes.

aluminium (Al)

☐

beryllium (Be)

☐

calcium (Ca)

☐

copper (Cu)

☐

francium (Fr)

☐

rubidium (Rb)

☐

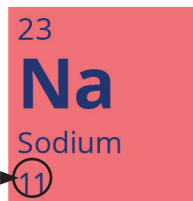
The atomic number increases as you move from left to right across a period.

4. What is the atomic number of the elements below?

a. molybdenum (Mo) \_\_\_\_\_

b. germanium (Ge) \_\_\_\_\_

atomic  
number



The atomic radius decreases as you move from left to right across a period. Some elements are shown in the tables below. The atomic radius of each element is given.

5. Complete the tables to predict the atomic radius of nitrogen and sulfur.

Element	boron	carbon	nitrogen	oxygen
Atomic Radius	87	67		48

Element	phosphorus	sulfur	chlorine	argon
Atomic Radius	98		79	71