



Groups and Periods on the Periodic Table

1. Write the group numbers in the boxes along the top of the periodic table.
2. Write the period numbers in the boxes down the side of the periodic table.

[illegible]



Elements in the same group have similar properties.

Lithium (Li) and potassium (K) are metals. They are solids at room temperature, have low melting points compared to other metals and are very reactive.

3. Name two other elements that have similar properties to lithium and potassium.

1. _____

2. _____

Neon (Ne) and xenon (Xe) are non-metals. They are gases at room temperature and are unreactive.

4. Name two other elements that have similar properties to neon and xenon.

1. _____

2. _____

The atomic number increases as you move from left to right across a period.

5. What is the atomic number of the elements below?

a. molybdenum (Mo) _____

b. germanium (Ge) _____

The atomic radius decreases as you move from left to right across a period. Some elements are shown in the table below. The atomic radius of each element is given in the table.

Element	Atomic Radius (pm)
argon	71
boron	87
carbon	67
chlorine	79
oxygen	48
phosphorus	98

6. Predict the atomic radius of the elements below.

a. nitrogen (N) _____ pm

b. sulfur (S) _____ pm