



Chemical Changes

Multiple Choice Questions **Answer Sheet**

Set 1

1 - C
2 - C
3 - C
4 - B
5 - A
6 - D
7 - B
8 - B
9 - C
10 - A

Set 2

1 - C
2 - D
3 - C
4 - A
5 - D
6 - D
7 - B
8 - A
9 - C
10 - D

Set 3

1 - A
2 - B
3 - A
4 - C
5 - A
6 - B
7 - B
8 - C
9 - C
10 - D

Set 4

1 - C
2 - B
3 - D
4 - A
5 - D
6 - B
7 - B
8 - D
9 - B
10 - A

Set 5

1 - D
2 - B
3 - D
4 - B
5 - C
6 - A
7 - B
8 - A
9 - A
10 - C



Chemical Changes

Multiple Choice Questions **Answer Sheet**

Set 1

1 - C
2 - C
3 - C
4 - B
5 - A
6 - D
7 - B
8 - B
9 - C
10 - A

Set 2

1 - C
2 - D
3 - C
4 - A
5 - D
6 - D
7 - B
8 - A
9 - C
10 - D

Set 3

1 - A
2 - B
3 - A
4 - C
5 - A
6 - B
7 - B
8 - C
9 - C
10 - D

Set 4

1 - C
2 - B
3 - D
4 - A
5 - D
6 - B
7 - B
8 - D
9 - B
10 - A

Set 5

1 - D
2 - B
3 - D
4 - B
5 - C
6 - A
7 - B
8 - A
9 - A
10 - C

Working for Calculations

Set 5

7. number of moles = concentration \times volume

$$\text{number of moles} = 0.25 \times 5 = \mathbf{1.25}$$

8. concentration = $\frac{\text{number of moles}}{\text{volume}}$

$$\text{concentration} = \frac{0.5}{4} = \mathbf{0.125\text{mol/dm}^3}$$

9. $\text{KOH} + \text{HCl} \longrightarrow \text{KCl} + \text{H}_2\text{O}$

$$\frac{30}{1000} = 0.03\text{dm}^3$$

$$\text{number of moles of KOH} = 0.5 \times 0.03 = 0.015$$

$$\text{so number of moles of HCl} = 0.015$$

$$\frac{60}{1000} = 0.06\text{dm}^3$$

$$\text{concentration of HCl} = \frac{0.015}{0.06} = \mathbf{0.25\text{mol/dm}^3}$$

10. $\text{NaOH} + \text{HCl} \longrightarrow \text{NaCl} + \text{H}_2\text{O}$

$$\frac{50}{1000} = 0.05\text{dm}^3$$

$$\text{number of moles of HCl} = 0.50 \times 0.05 = 0.025$$

$$\text{so number of moles of NaOH} = 0.025$$

$$\text{volume of NaOH} = \frac{0.025}{1.25} = 0.02\text{dm}^3$$

$$0.02 \times 1000 = \mathbf{20\text{cm}^3}$$